



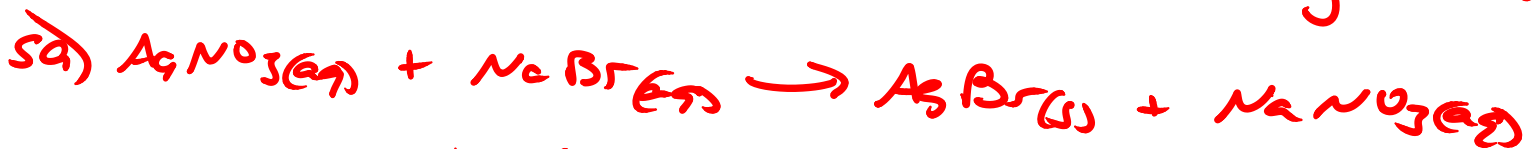
$$0.0750 \text{ L} \frac{0.100 \text{ mole AgNO}_3}{1 \text{ L}} \Bigg| \frac{1 \text{ mole Na}_2\text{CrO}_4}{2 \text{ mole AgNO}_3} \Bigg| \frac{161.98 \text{ g Na}_2\text{CrO}_4}{1 \text{ mole Na}_2\text{CrO}_4} = 0.607 \text{ g Na}_2\text{CrO}_4$$



$$0.0500 \text{ L} \frac{0.200 \text{ mole}}{1 \text{ L}} = 0.0100 \text{ mole Al}(\text{NO}_3)_3$$

$$0.2000 \text{ L} \frac{0.100 \text{ mole}}{1 \text{ L}} = 0.0200 \text{ mole KOH} \quad \text{limit}$$

$$0.0200 \text{ mole KOH} \frac{1 \text{ mole Al}(\text{OH})_3}{3 \text{ mole KOH}} \Bigg| \frac{78.00 \text{ g Al}(\text{OH})_3}{1 \text{ mole Al}(\text{OH})_3} = 0.520 \text{ g Al}(\text{OH})_3$$



$$100.0 \text{ mL AgNO}_3 \frac{1 \text{ L}}{1000 \text{ mL}} \Bigg| \frac{0.150 \text{ mole AgNO}_3}{1 \text{ L AgNO}_3} = 1.50 \times 10^{-2} \text{ mole AgNO}_3$$

$$20.0 \text{ mL NaBr} \frac{1 \text{ L}}{1000 \text{ mL}} \Bigg| \frac{1.00 \text{ mole NaBr}}{1 \text{ L}} = 2.00 \times 10^{-2} \text{ mole NaBr}$$

$$1.50 \times 10^{-2} \text{ mole AgNO}_3 \frac{1 \text{ mole AgBr}}{1 \text{ mole AgNO}_3} \Bigg| \frac{187.8 \text{ g AgBr}}{1 \text{ mole AgBr}} = 2.82 \text{ g AgBr}$$



$$0.1000\text{L KOH} \frac{0.200\text{mole KOH}}{1\text{L}} = \frac{2.00 \times 10^{-2}}{2} \text{ KOH.} \quad \# \text{ limits}$$

$$0.1000\text{L Mg}(\text{NO}_3)_2 \frac{0.200\text{mole Mg}(\text{NO}_3)_2}{1\text{L Mg}(\text{NO}_3)_2} = \frac{2.00 \times 10^{-2}\text{mole}}{1} \text{ Mg}(\text{NO}_3)_2$$

$$0.0200\text{mole KOH} \frac{1\text{mole Mg}(\text{OH})_2}{2\text{mole KOH}} \left| \frac{58.33\text{g Mg}(\text{OH})_2}{1\text{mole Mg}(\text{OH})_2} \right. = 0.583\text{g Mg}(\text{OH})_2$$



I	$2.00 \times 10^{-2}\text{mole}$	$2.00 \times 10^{-2}\text{mole}$	0
L	-1.00×10^{-2}	-2.00×10^{-2}	$+1.00 \times 10^{-2}\text{mole}$
F	$1.00 \times 10^{-2}\text{mole}$	0	$1.00 \times 10^{-2}\text{mole}$

$$[\text{OH}^{-}] = 0$$

$$[\text{Mg}^{+2}] = \frac{1.00 \times 10^{-2}\text{mole}}{0.1000\text{L} + 0.1000\text{L}} = 5.00 \times 10^{-2}\text{M}$$

$$[\text{K}^{+}] = \frac{0.0200\text{mole}}{0.2000\text{L}} = 0.100\text{M}$$

$$[\text{NO}_3^{-}] = \frac{0.0400\text{mole}}{0.2000\text{L}} = 0.200\text{M}$$