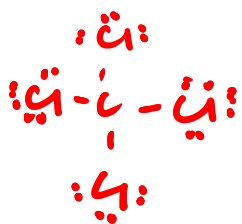


103) a) CCl_4



tetrahedral
 109.5°

b) PCl_3



trigonal pyramidal
 $< 109.5^\circ$

c) $SeCl_2$



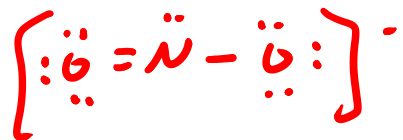
bent
 $< 109.5^\circ$

d) ICl

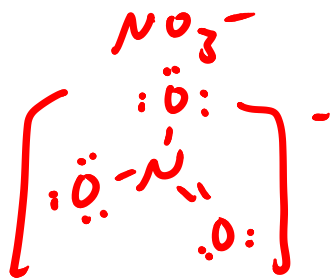


linear
no bond angle

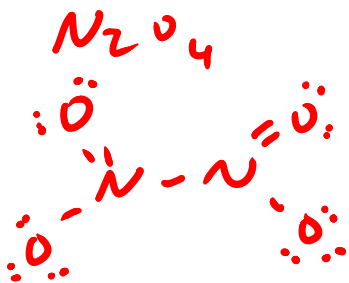
87 a) NO_2^-



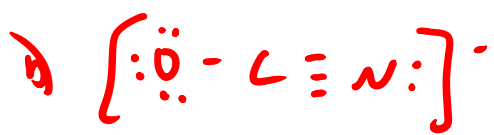
bent
 $\hat{=} 120^\circ$



trigonal planar
 120°



trigonal planar
 120°



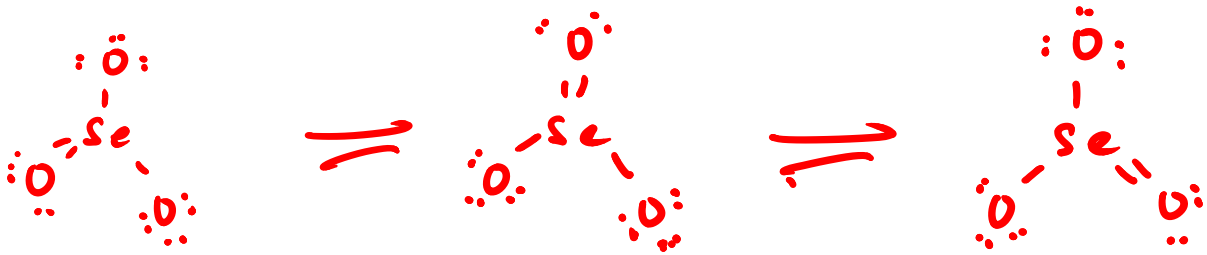
all linear



180°



107) a) SeO_3 (24e⁻)



trigonal planar 120°

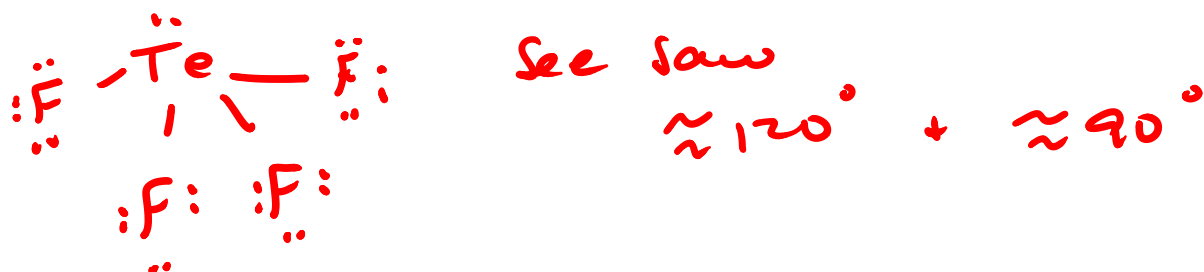
b) SeO_2 (18e⁻)

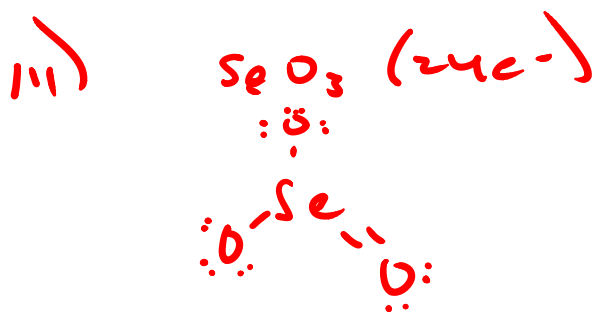


bent
 $\approx 120^\circ$

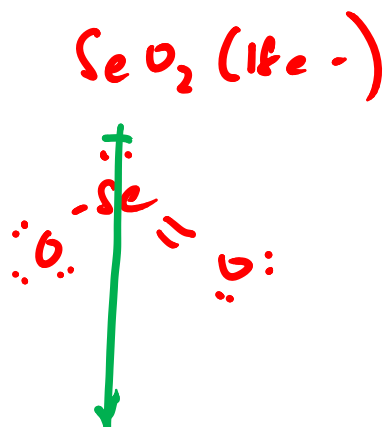
109 a) XeU_2 (22e⁻)



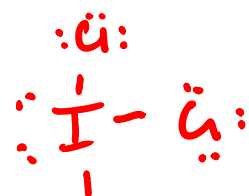




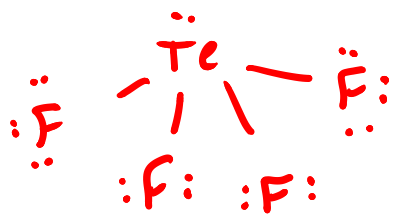
bonds are polar
but molecule is
non polar



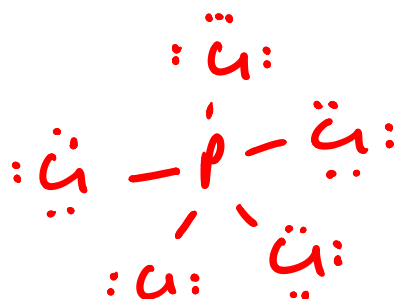
bond dipoles cancel
no net dipole



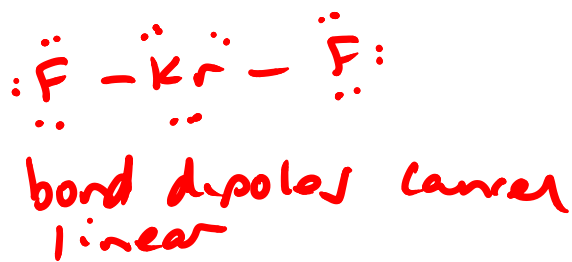
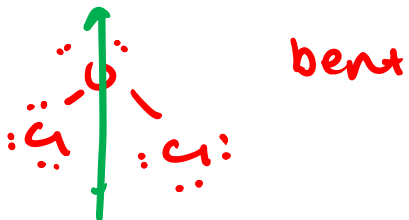
net dipole



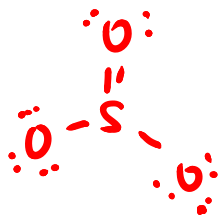
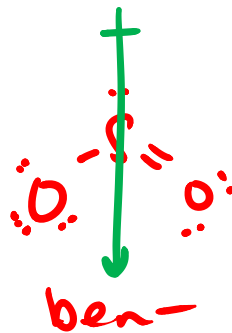
net dipole



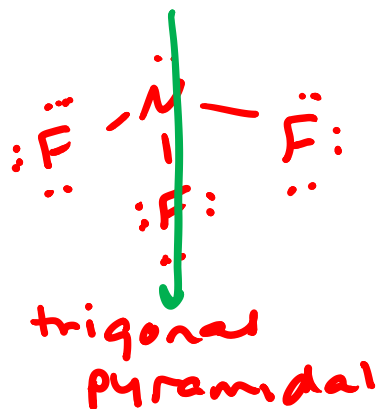
bond dipoles cancel
no net dipole

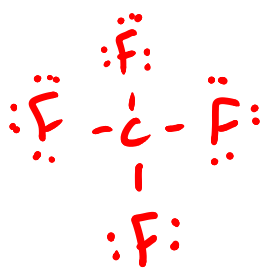


bond dipoles cancel
linear

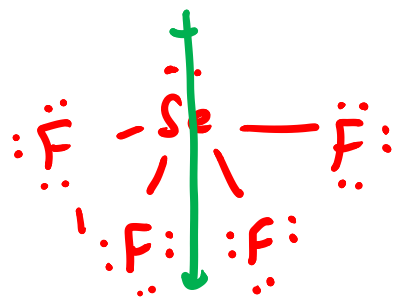


bond dipoles cancel
trigonal planar

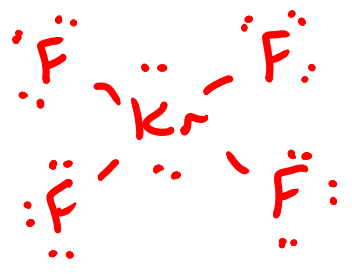




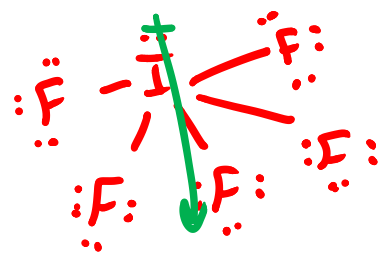
bond dipoles
cancel
tetrahedral



see saw



square planar
bond dipoles cancel



square pyramidal

