**LAB TYPES OF REACTIONS**

This activity will allow you to experiment with each of the five types of chemical reactions. You will follow the procedures given below and then write a balanced chemical equation for each reaction. No formal lab report will be due for this lab. Make sure to include detailed observations and a balanced chemical equation for each experiment.

**Section 1: Synthesis Reaction**

1. Obtain a small piece of magnesium ribbon from your teacher.
2. Light a Bunsen Burner, making sure to adjust the flame so a bright blue cone is visible
3. Using crucible tongs, hold the magnesium ribbon directly over the blue cone of the flame.
4. Make observations and a balanced equation.
5. Clean up your area.

**Section 2: Decomposition Reaction**

1. Place 10 ml of peroxide in a test tube
2. Add a small piece of potato to the peroxide. There is a chemical in the potato that will increase the rate of decomposition of peroxide.
3. Make observations and a balanced equation.
4. Clean up your area.

**Section 3: Single Displacement Reaction**

1. Place 10 ml of sulfuric acid (hydrogen sulfate) in a flask.
2. Add a small amount of zinc metal into the sulfuric acid.
3. Make observations and a balanced equation.
4. Clean up your area.

**Section 4: Double Displacement Reaction**

1. On a watch glass, place 5 drops of potassium iodide solution.
2. \*\*\*Do Not Mix The Droppers up!!!
3. To the potassium iodide, add 5 drops of lead II nitrate.
4. Make observations and a balanced equation.
5. Clean up your area.

**Section 5: Combustion Reaction**

1. Add 5 drops of isopropanol (C3H7OH) to a watch glass.
2. Carefully light the isopropanol with a match, using tongs to keep your hands away from the flame.
3. Make observations and a balanced equation.
4. Clean up your area.

**Lab**

**Reaction Types**

**Section 1: Synthesis**

Observations:

1.

2.

3.

Balanced Equation:

**Section 2: Decomposition**

Observations:

1.

2.

3.

Balanced Equation:

**Section 3: Single Displacement**

Observations:

1.

2.

3.

Balanced Equation:

**Section 4: Double Displacement**

Observations:

1.

2.

3.

Balanced Equation:

**Section 5: Combustion**

Observations:

1.

2.

3.

Balanced Equation: