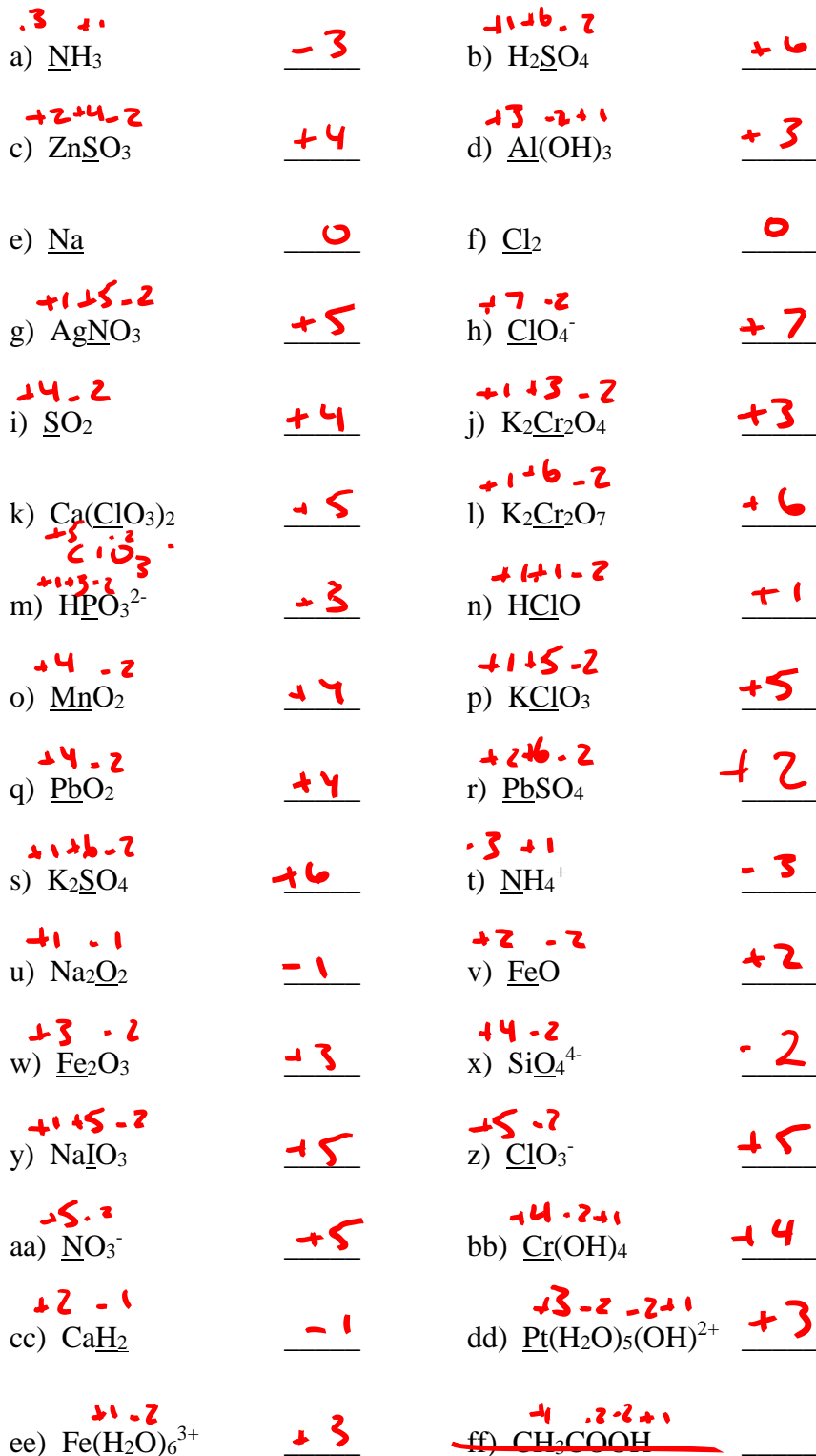


Name _____ Date _____

REVIEW REDOX

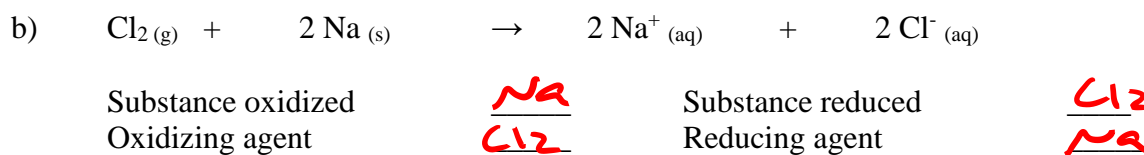
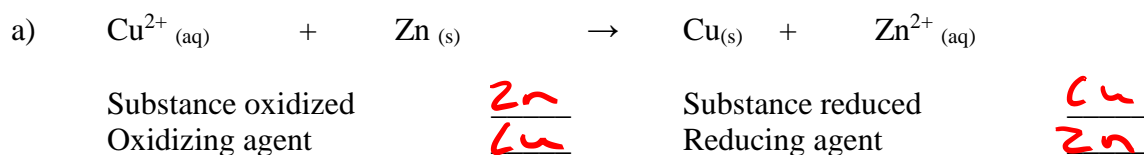
1. State the Oxidation Number of each of the elements that is underlined.



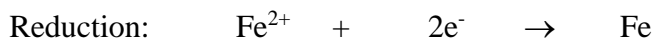
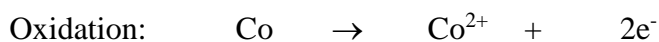
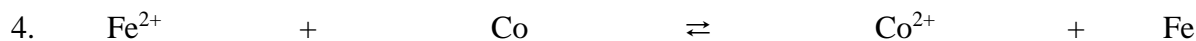
2. What is the oxidation number of carbon in each of the following substances?



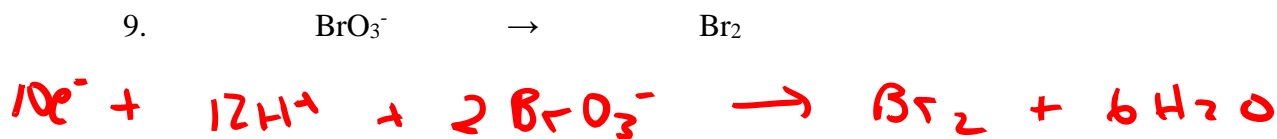
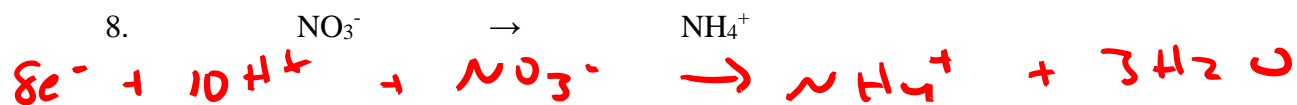
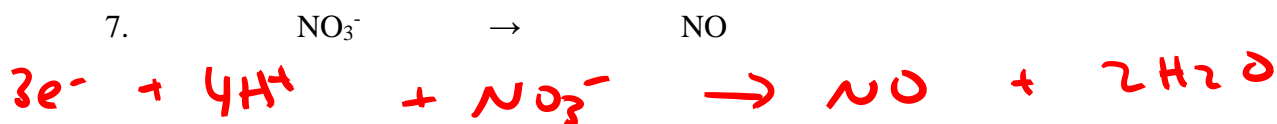
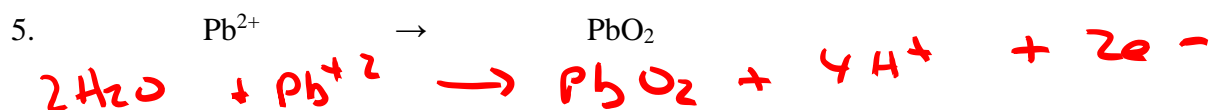
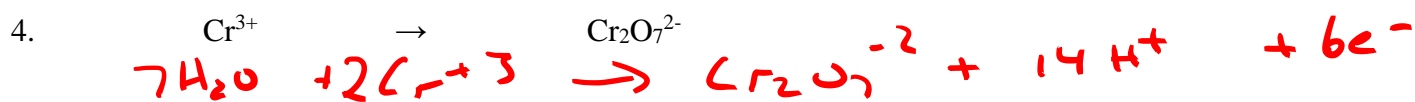
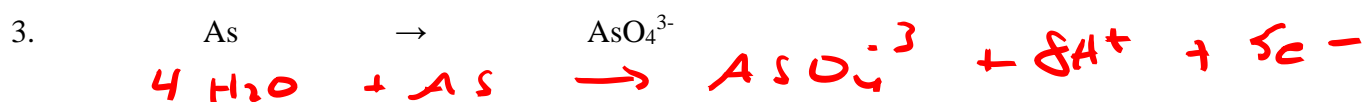
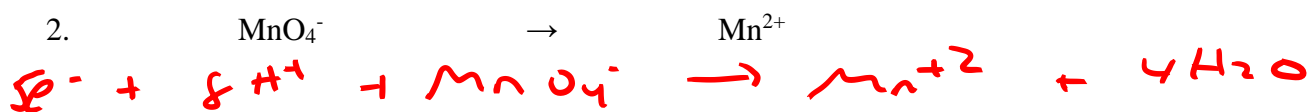
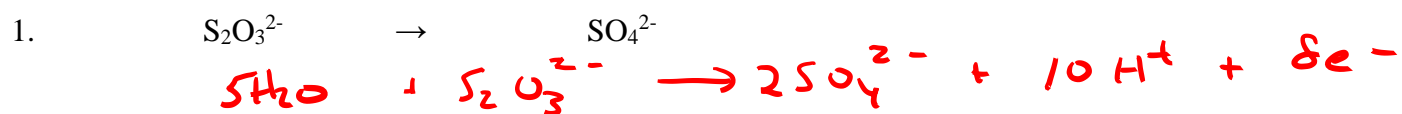
3. For each of the following reactants, identify: the oxidizing agent, the reducing agent, the substance oxidized and the substance reduced.



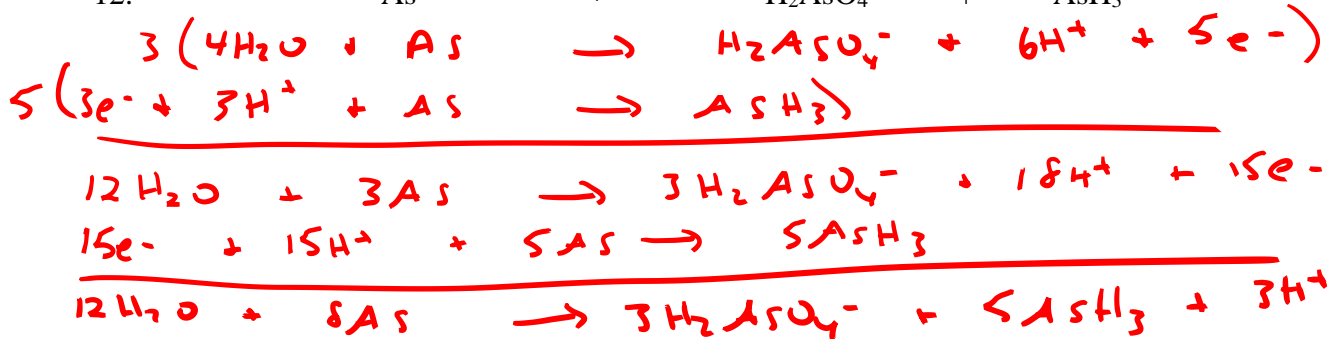
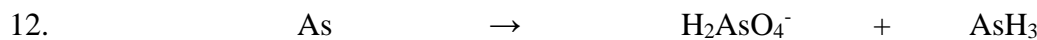
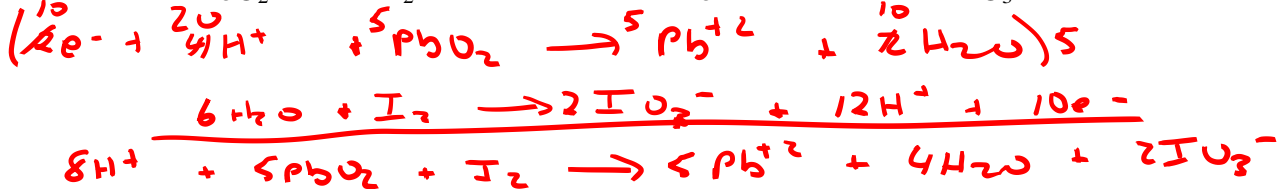
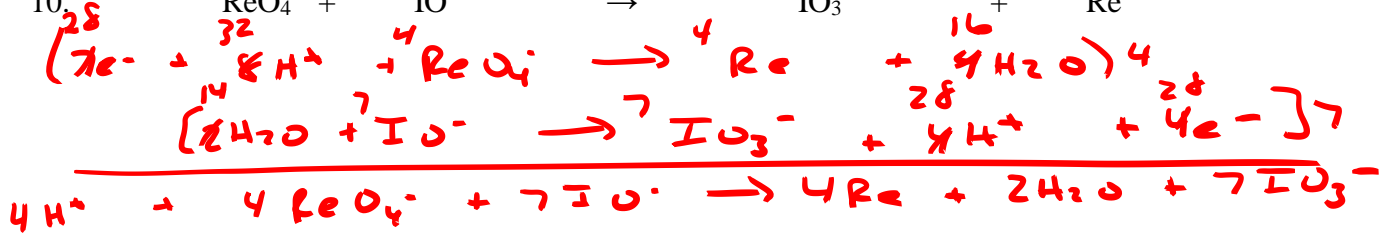
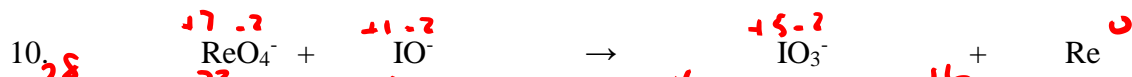
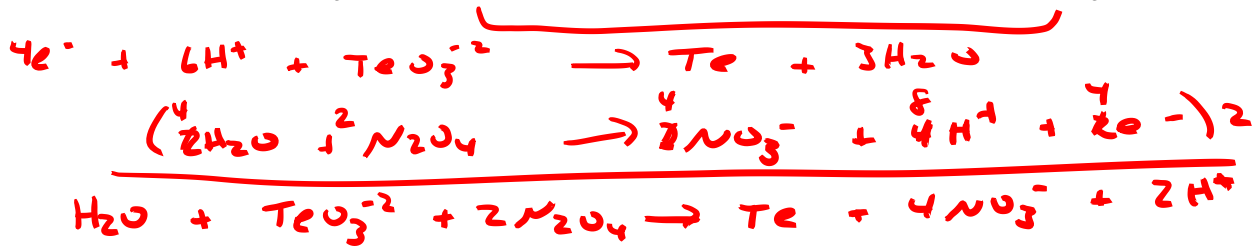
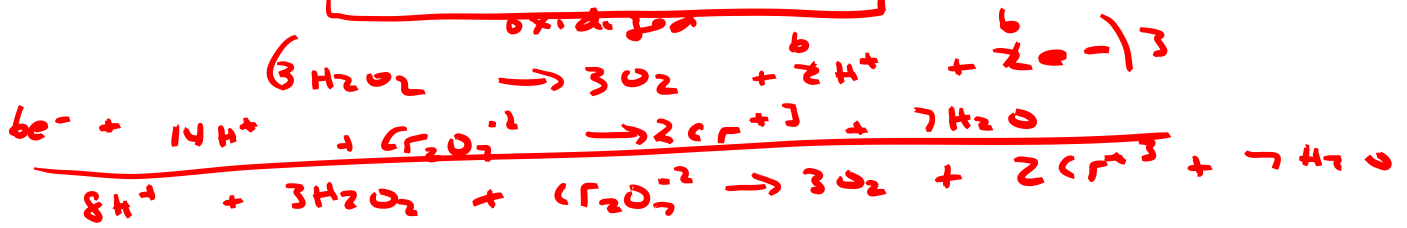
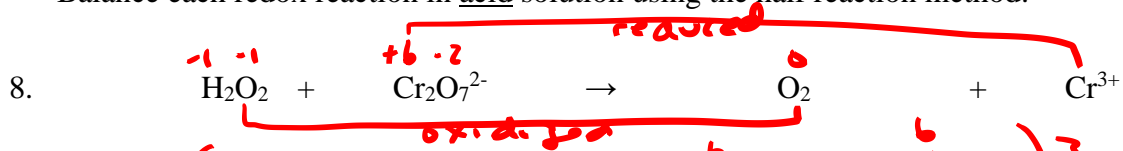
Write the oxidation and reduction reactions for each redox reaction. The first one is done for you.



Balance each of the following half-cell reactions. (In each case assume that the reaction takes place in an **ACIDIC** solution.) Also, state whether the reaction is oxidation or reduction.



Balance each redox reaction in acid solution using the half reaction method.



Balance each redox reaction in basic solution using the half reaction method.

